

Nuclide ^[3]_[n 1]	Historic name	Z	N	Isotopic mass (Da) ^[4]_{[n 2][n 3]}	Half-life	Decay mode _[n 4]	Daughter isotope _{[n 5][n 6]}	Spin and parity _{[n 7][n 8]}	Natural abundance (mole fraction)	
		Excitation energy ^[n 8]							Normal proportion	Range of variation
²¹⁴ U ^[5]		92	122		0.52(+0.95−0.21) ms	α	²¹⁰ Th	0+		
²¹⁵ U ^[6]		92	123	215.026760(90)	2.24 ms	α	²¹¹ Th	5/2−#		
²¹⁶ U ^[6]^[7]		92	124	216.024760(30)	2.25(+0.63−0.40) ms ^[5]	α	²¹² Th	0+		
^{216m} U ^[8]					1.31 ms			8+		
²¹⁷ U		92	125	217.02437(9)	²⁶ (14) ms [16(+21−6) ms]	α	²¹³ Th	1/2−#		
²¹⁸ U		92	126	218.02354(3)	0.65(+0.08−0.07) ms ^[5]	α	²¹⁴ Th	0+		
²¹⁹ U		92	127	219.02492(6)	⁵⁵ (25) μs [42(+34−13) μs]	α	²¹⁵ Th	9/2+#		
²²¹ U ^[9]		92	129	221.02640(11)#	0.66(14) μs	α	²¹⁷ Th	(9/2+)		
²²² U		92	130	222.02609(11)#	1.4(7) μs [1.0(+10−4) μs]	α β ⁺ (10 ^{−6} %)	²¹⁸ Th ²²² Pa	0+		
²²³ U		92	131	223.02774(8)	²¹ (8) μs [18(+10−5) μs]	α	²¹⁹ Th	7/2+#		
²²⁴ U		92	132	224.027605(27)	940(270) μs	α	²²⁰ Th	0+		
²²⁵ U		92	133	225.02939#	61(4) ms	α	²²¹ Th	(5/2+) #		
²²⁶ U		92	134	226.029339(14)	269(6) ms	α	²²² Th	0+		
²²⁷ U		92	135	227.031156(18)	1.1(1) min	α β ⁺ (.001%)	²²³ Th ²²⁷ Pa	(3/2+)		
²²⁸ U		92	136	228.031374(16)	9.1(2) min	α (95%) EC (5%)	²²⁴ Th ²²⁸ Pa	0+		
²²⁹ U		92	137	229.033506(6)	58(3) min	β ⁺ (80%) α (20%)	²²⁹ Pa ²²⁵ Th	(3/2+)		
²³⁰ U		92	138	230.033940(5)	20.8 d	α SF (1.4×10 ^{−10} %) β ⁺ β ⁺ (rare)	²²⁶ Th (various) ²³⁰ Th	0+		
²³¹ U		92	139	231.036294(3)	4.2(1) d	EC α (.004%)	²³¹ Pa ²²⁷ Th	(5/2)(+ #)		
²³² U		92	140	232.0371562(24)	68.9(4) y	α CD (8.9×10 ^{−10} %) CD (5×10 ^{−12} %) SF (10 ^{−12} %)	²²⁸ Th ²⁰⁸ Pb ²⁴ Ne ²⁰⁴ Hg ²⁸ Mg (various)	0+		
²³³ U		92	141	233.0396352(29)	1.592(2)×10 ⁵ y	α SF (6×10 ^{−9} %) CD (7.2×10 ^{−11} %) CD (1.3×10 ^{−13} %)	²²⁹ Th (various) ²⁰⁹ Pb ²⁴ Ne ²⁰⁵ Hg ²⁸ Mg	5/2+	Trace ^[n 9]	
²³⁴ U ^{[n 10][n 11]}	Uranium II	92	142	234.0409521(20)	2.455(6)×10 ⁵ y	α SF (1.73×10 ^{−9} %) CD (1.4×10 ^{−11} %) CD (9×10 ^{−12} %)	²³⁰ Th (various) ²⁰⁶ Hg ²⁸ Mg ¹⁸⁴ Hf ²⁶ Ne ²⁴ Ne	0+	[0.000054(5)] ^[n 12]	0.000050–0.000059
^{234m} U				1421.32(10) keV	33.5(20) ms			6−		
²³⁵ U ^{[n 13][n 14][n 15]}	Actino Uranium Actino-Uranium	92	143	235.0439299(20)	7.038(1)×10 ⁸ y	α SF (7×10 ^{−9} %) CD (8×10 ^{−10} %)	²³¹ Th (various) ¹⁸⁶ Hf ²⁵ Ne ²⁴ Ne	7/2−	[0.007204(6)]	0.007198–0.007207
^{235m} U				0.0765(4) keV	~26 min	IT	²³⁵ U	1/2+		
²³⁶ U	Thoruranium ^[10]	92	144	236.045568(2)	2.342(3)×10 ⁷ y	α SF (9.6×10 ^{−8} %)	²³² Th (various)	0+	Trace ^[n 16]	
^{236m1} U				1052.89(19) keV	100(4) ns			(4)−		
^{236m2} U				2750(10) keV	120(2) ns			(0+)		
²³⁷ U		92	145	237.0487302(20)	6.75(1) d	β [−]	²³⁷ Np	1/2+	Trace ^[n 17]	
²³⁸ U ^{[n 11][n 13][n 14]}	Uranium I	92	146	238.0507882(20)	4.468(3)×10 ⁹ y	α SF (5.45×10 ^{−5} %) β [−] β [−] (2.19×10 ^{−10} %)	²³⁴ Th (various) ²³⁸ Pu	0+	[0.992742(10)]	0.992739–0.992752

^{238m} U			2557.9(5) keV	280(6) ns		0+		
²³⁹ U		92	147	239.0542933(21)	23.45(2) min	β [−]	²³⁹ Np	5/2+
^{239m1} U				20(20)# keV	>250 ns			(5/2+)
^{239m2} U				133.7990(10) keV	780(40) ns			1/2+
²⁴⁰ U		92	148	240.056592(6)	14.1(1) h	β [−] α (10 ^{−10} %)	²⁴⁰ Np ²³⁶ Th	0+ Trace ^[n 18]
²⁴² U		92	150	242.06293(22)#	16.8(5) min	β [−]	²⁴² Np	0+

- ^mU – Excited **nuclear isomer**.
- () – Uncertainty (1σ) is given in concise form in parentheses after the corresponding last digits.
- # – Atomic mass marked #: value and uncertainty derived not from purely experimental data, but at least partly from trends from the Mass Surface (TMS).
- Modes of decay:
 - CD: **Cluster decay**
 - EC: **Electron capture**
 - SF: **Spontaneous fission**
- Bold italics symbol*** as daughter – Daughter product is nearly stable.
- Bold symbol** as daughter – Daughter product is stable.
- () spin value – Indicates spin with weak assignment arguments.

- # – Values marked # are not purely derived from experimental data, but at least partly from trends of neighboring nuclides (TNN).
- Intermediate decay product of ²³⁷Np
- Used in **uranium–thorium dating**
- Used in **uranium–uranium dating**
- Intermediate **decay product** of ²³⁸U
- Primordial radionuclide**
- Used in **Uranium–lead dating**
- Important in nuclear reactors
- Intermediate decay product of ²⁴⁴Pu, also produced by **neutron capture** of ²³⁵U
- Neutron capture product, parent of trace quantities of ²³⁷Np
- Intermediate decay product of ²⁴⁴Pu

Actinides vs fission products

Actinides and fission products by half-life								
Actinides ^[11] by decay chain				Half-life range (a)	Fission products of ²³⁵ U by yield ^[12]			
4n	4n+1	4n+2	4n+3		4.5–7%	0.04–1.25%	<0.001%	
²²⁸ Ra ^{Ne}				4–6 a	†	¹⁵⁵ Eu ^b		
²⁴⁴ Cm ^f	²⁴¹ Pu ^f	²⁵⁰ Cf	²²⁷ Ac ^{Ne}	10–29 a		⁹⁰ Sr	⁸⁵ Kr	
²³² U ^f		²³⁸ Pu ^f	²⁴³ Cm ^f	29–97 a		¹³⁷ Cs	¹⁵¹ Sm ^b	
²⁴⁸ Bk ^[13]	²⁴⁹ Cf ^f	^{242m} Am ^f		141–351 a			^{121m} Sn	
	²⁴¹ Am ^f		²⁵¹ Cf ^f ^[14]	430–900 a				
		²²⁶ Ra ^{Ne}	²⁴⁷ Bk	1.3–1.6 ka	No fission products have a half-life in the range of 100–210 ka ...			
²⁴⁰ Pu	²²⁹ Th	²⁴⁶ Cm ^f	²⁴³ Am ^f	4.7–7.4 ka				
		²⁴⁵ Cm ^f	²⁵⁰ Cm	8.3–8.5 ka				
			²³⁹ Pu ^f	24.1 ka				
		²³⁰ Th ^{Ne}	²³¹ Pa ^{Ne}	32–76 ka				
²³⁶ Np ^f	²³³ U ^f	²³⁴ U ^{Ne}		150–250 ka	‡	⁹⁹ Tc [℄]	¹²⁸ Sn	
²⁴⁸ Cm		²⁴² Pu		327–375 ka		⁷⁹ Se [℄]		
				1.53 Ma		⁹³ Zr		
	²³⁷ Np ^f			2.1–6.5 Ma		¹³⁵ Cs [℄]	¹⁰⁷ Pd	
²³⁶ U			²⁴⁷ Cm ^f	15–24 Ma		¹²⁹ I [℄]		
²⁴⁴ Pu				80 Ma				
²³² Th ^{Ne}		²³⁸ U ^{Ne}	²³⁵ U ^f ^{Ne}	0.7–14.1 Ga			... nor beyond 15.7 Ma ^[15]	

Legend for superscript symbols
 ℄ has thermal **neutron capture** cross section in the range of 8–50 barns
 f fissile
 m **metastable isomer**
 Ne **primarily a naturally occurring radioactive material (NORM)**
 b **neutron poison** (thermal neutron capture cross section greater than 3k barns)
 † range 4–97 a: **Medium-lived fission product**
 ‡ over 200 ka: **Long-lived fission product**

Uranium-232

Uranium-232 has a half-life of 68.9 years and is a side product in the thorium cycle. It has been cited as an obstacle to nuclear proliferation using ²³³U as the fissile material, because the intense **gamma radiation** emitted by ²⁰⁸Tl (a daughter of ²³²U, produced relatively quickly) makes the ²³³U contaminated with it more difficult to handle. Uranium-232 is a rare example of an **even-even isotope** that is **fissile** with both thermal and fast neutrons.^{[16][17]}

Uranium-233

Uranium-233 is a fissile isotope of uranium that is bred from **thorium-232** as part of the thorium fuel cycle. Uranium-233 was investigated for use in nuclear weapons and as a reactor fuel; however, it was never deployed in nuclear weapons or used commercially as a nuclear fuel.^[18] It has been used successfully in experimental nuclear reactors and has been proposed for much wider use as a nuclear fuel. It has a half-life of 1.592 × 10⁵ years.

Uranium-233 is produced by the neutron irradiation of thorium-232. When thorium-232 absorbs a **neutron**, it becomes thorium-233, which has a half-life of only 22 minutes. Thorium-233 decays into **protactinium-233** through **beta decay**. Protactinium-233 has a half-life of 27 days and beta decays into uranium-233; some proposed molten salt reactor designs attempt to physically isolate the protactinium from further neutron capture before beta decay can occur.

Uranium-233 usually fissions on neutron absorption but sometimes retains the neutron, becoming **uranium-234**. The capture-to-fission ratio is smaller than the other two major fissile fuels **uranium-235** and **plutonium-239**; it is also lower than that of short-lived **plutonium-241**, but bested by very difficult-to-produce **neptunium-236**.

Uranium-234

Uranium-234 is an isotope of uranium found in uranium ore, ²³⁴U occurs as an indirect decay product of uranium-238, but it makes up only 0.0055% (5 parts per million) of the raw uranium because its half-life of just 245,500 years is only about 1/18,000 as long as that of ²³⁸U. The path of production of ²³⁴U via nuclear decay is as follows: ²³⁸U nuclei emit an alpha particle to become thorium-234. Next, with a short half-life, a ²³⁴Th nucleus emits a beta particle to become protactinium-234. Finally, ²³⁴Pa nuclei each emit another beta particle to become ²³⁴U nuclei.^{[19][20]}

²³⁴U nuclei usually last for hundreds of thousands of years, but then they decay by alpha emission to thorium-230, except for the small percentage of nuclei that undergo spontaneous fission.

Extraction of rather small amounts of ²³⁴U from natural uranium would be feasible using isotope separation, similar to that used for regular uranium-enrichment. However, there is no real demand in chemistry, physics, or engineering for isolating ²³⁴U. Very small pure samples of ²³⁴U can be extracted via the chemical ion-exchange process—from samples of plutonium-238 that have been aged somewhat to allow some decay to ²³⁴U via alpha emission.

Enriched uranium contains more ²³⁴U than natural uranium as a byproduct of the uranium enrichment process aimed at obtaining uranium-235, which concentrates lighter isotopes even more strongly than it does ²³⁵U. The increased percentage of ²³⁴U in enriched natural uranium is acceptable in current nuclear reactors, but (re-enriched) reprocessed uranium might contain even higher fractions of ²³⁴U, which is undesirable.^[21] This is because ²³⁴U is not fissile, and tends to absorb slow neutrons in a nuclear reactor—becoming ²³⁵U.^{[20][21]}

²³⁴U has a neutron capture cross section of about 100 barns for thermal neutrons, and about 700 barns for its resonance integral—the average over neutrons having various intermediate energies. In a nuclear reactor, non-fissile isotopes capture a neutron breeding fissile isotopes. ²³⁴U is converted to ²³⁵U more easily and therefore at a greater rate than uranium-238 is to plutonium-239 (via neptunium-239), because ²³⁸U has a much smaller neutron-capture cross section of just 2.7 barns.

Uranium-235

Uranium-235 is an isotope of uranium making up about 0.72% of natural uranium. Unlike the predominant isotope uranium-238, it is fissile, i.e., it can sustain a fission chain reaction. It is the only fissile isotope that is a primordial nuclide or found in significant quantity in nature.

Uranium-235 has a half-life of 703.8 million years. It was discovered in 1935 by Arthur Jeffrey Dempster. Its (fission) nuclear cross section for slow thermal neutron is about 504.81 barns. For fast neutrons it is on the order of 1 barn. At thermal energy levels, about 5 of 6 neutron absorptions result in fission and 1 of 6 result in neutron capture forming uranium-236.^[22] The fission-to-capture ratio improves for faster neutrons.

Uranium-236

Uranium-236 is an isotope of uranium that is neither fissile with thermal neutrons, nor very good fertile material, but is generally considered a nuisance and long-lived radioactive waste. It is found in spent nuclear fuel and in the reprocessed uranium made from spent nuclear fuel.

Uranium-237

Uranium-237 is an isotope of uranium. It has a half life of about 6.75(1) days. It decays into neptunium-237 by beta decay.

Uranium-238

Uranium-238 (²³⁸U or U-238) is the most common isotope of uranium found in nature. It is not fissile, but is a fertile material: it can capture a slow neutron and after two beta decays become fissile plutonium-239. Uranium-238 is fissionable by fast neutrons, but cannot support a chain reaction because inelastic scattering reduces neutron energy below the range where fast fission of one or more next-generation nuclei is probable. Doppler broadening of ²³⁸U's neutron absorption resonances, increasing absorption as fuel temperature increases, is also an essential negative feedback mechanism for reactor control.

Around 99.284% of natural uranium is uranium-238, which has a half-life of 1.41×10¹⁷ seconds (4.468×10⁹ years, or 4.468 billion years). Depleted uranium has an even higher concentration of the ²³⁸U isotope, and even low-enriched uranium (LEU), while having a higher proportion of the uranium-235 isotope (in comparison to depleted uranium), is still mostly ²³⁸U. Reprocessed uranium is also mainly ²³⁸U, with about as much uranium-235 as natural uranium, a comparable proportion of uranium-236, and much smaller amounts of other isotopes of uranium such as uranium-234, uranium-233, and uranium-232.

Uranium-239

Uranium-239 is an isotope of uranium. It is usually produced by exposing ²³⁸U to neutron radiation in a nuclear reactor. ²³⁹U has a half-life of about 23.45 minutes and decays into neptunium-239 through beta decay, with a total decay energy of about 1.29 MeV.^[23] The most common gamma decay at 74.660 keV accounts for the difference in the two major channels of beta emission energy, at 1.28 and 1.21 MeV.^[24]

²³⁹Np further decays to plutonium-239 also through beta decay (²³⁹Np has a half-life of about 2.356 days), in a second important step that ultimately produces fissile ²³⁹Pu (used in weapons and for nuclear power), from ²³⁸U in reactors.

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11. Plus radium (element 88). While actually a sub-actinide, it immediately precedes actinium (89) and follows a three-element gap of instability after polonium (84) where no nuclides have half-lives of at least four years (the longest-lived nuclide in the gap is radon-222 with a half life of less than four days). Radium's longest lived isotope, at 1,600 years, thus merits the element's inclusion here.
12. Specifically from thermal neutron fission of U-235, e.g. in a typical nuclear reactor.
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"The isotopic analyses disclosed a species of mass 248 in constant abundance in three samples analysed over a period of about 10 months. This was ascribed to an isomer of Bk²⁴⁸ with a half-life greater than 9 [years]. No growth of Cf²⁴⁸ was detected, and a lower limit for the β⁻ half-life can be set at about 10⁴ [years]. No alpha activity attributable to the new isomer has been detected; the alpha half-life is probably greater than 300 [years]."
14. This is the heaviest nuclide with a half-life of at least four years before the "Sea of Instability".
15. Excluding those "classically stable" nuclides with half-lives significantly in excess of ²³²Th; e.g., while ^{113m}Cd has a half-life of only fourteen years, that of ¹¹³Cd is nearly eight quadrillion years.
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